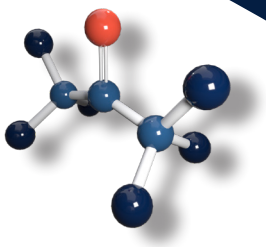
Gas Factsheet



Acetone

Formula: C_3H_6O
CAS: 67-64-1
Source: Volcanic gases, Exhausts fumes, Cigarette smoke & Landfill sites
Detection Method: Tiger, Tiger^{LT}, Cub, Falco, TVOC 2, GasClam 2 & GasCheck G

Acetone, a colourless liquid also known as Propanone, is a solvent used in manufacture of plastics and other industrial products. Acetone may also be used to a limited extent in household products, including cosmetics and personal care products, where its most frequent application would be in the formulation of nail polish removers. Acetone occurs naturally in the human body as a byproduct of metabolism.

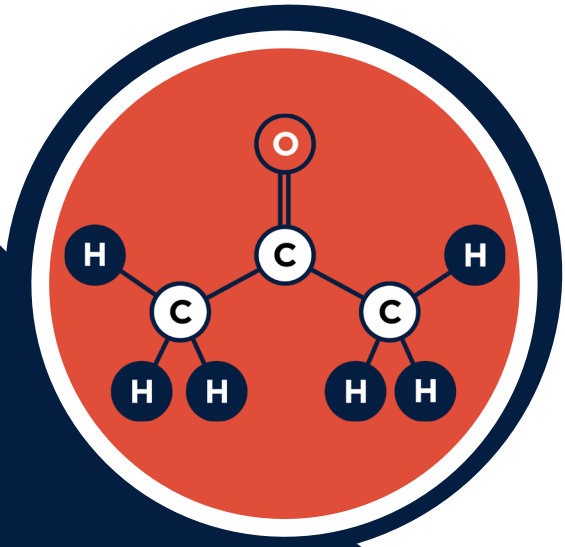
C_3H_6O - THE SMELL OF ACETONE ON SOMEONE'S BREATH IS A BAD SIGN

Why Do We Use Acetone?

Acetone is a solvent, which means it can break down or dissolve substances like paint and varnish. That's why it's an ingredient in nail polish removers, varnish removers, and paint removers. Companies also use this chemical to remove grease from wool, reduce the stickiness of silk, and make protective coatings for furniture and cars.

Acetone Within Our Bodies

Acetone is a chemical that is found naturally in the environment and is also produced by industries. Low levels of acetone are normally present in the body from the breakdown of fat; the body can use it in normal processes that make sugar and fat. The key thing to remember it that Acetone is only safe for our bodies when it is produced internally and released out into the environment. **If acetone goes in the opposite direction, from external sources into our bodies, then it will cause minor to severe health problems.**



Exposure To Acetone

If you are exposed to acetone, it goes into your blood which then carries it to all the organs in your body. If it is a small amount, the liver breaks it down to chemicals that are not harmful and uses these chemicals to make energy for normal body functions.

Exposure to acetone results mostly from breathing air, drinking water, or coming in contact with products or soil that contain acetone. Exposure to moderate-to-high amounts of acetone can irritate your eyes and respiratory system, and make you dizzy. Very high exposure may cause you to lose consciousness.

History Of Acetone

Acetone was first discovered during the middle Ages by alchemists, It was French chemist **Jean Baptiste Dumas** and German chemist **Justus von Liebig**, who determined acetone chemical formula. In 1833, it was named acetone, by adding one-suffix to the stem of the corresponding acetic acid. By 1852, **Alexander William Williamson**, an English chemist concluded that acetone is a methyl acetate and in the same year it was agreed/concurred by a French chemist called **Charles Frederic Gerhardt**.

Acetone Detection Instruments

Fixed Instruments


Semi-Portable Instruments


Portable Instruments


Personal Instruments

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